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PHYSICAL CHEMISTRY: A Laboratory Manual has been designed to meet the need of graduate and postgraduate students. The language is simple and the students can perform the experiments themselves without much help from the teacher. In each chapter, complete theory has been introduced before the start of experiment.

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PHYSICAL CHEMISTRY By: Shailendra Kumar Page No.: 2 3. FACTORS AFFECTING RATE OF REACTION : (a) Nature of reaction (b) Temperature (c) Concentration of reactants (d) Catalyst (e) Surface area of solid reactant (f) Presence of light There are certain other factors like light, electrical field etc. also for certain particular reactions.

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PHYSICAL CHEMISTRY By: Shailendra Kumar 1. Opp. Khuda Baksh Library, Ashok Rajpath, Patna - 4 2. House no. 5A/65, Opp. Mahual Kothi, Alpana Market, Patna Page No.: 3 09. The complex ion, $[\text{Cr}(\text{NH}_3)_5\text{Cl}]^{2+}$, reacts with OH^- ion in aqueous solution, $[\text{Cr}(\text{NH}_3)_5\text{Cl}]^{2+}(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow [\text{Cr}(\text{NH}_3)_5(\text{OH})]^{2+}(\text{aq}) + \text{Cl}^-(\text{aq})$

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PHYSICAL CHEMISTRY By: Shailendra Kumar 1. Opp. Khuda Baksh Library, Ashok Rajpath, Patna - 4 2. House no. 5A/65, Opp. Mahual Kothi, Alpana Market, Patna Page No.: 1 1. Consider the following gas-phase reaction, at equilibrium, under conditions where the value of the equilibrium constant, K, is equal to 8: $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$, $K = 8$

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PHYSICAL CHEMISTRY By: Shailendra Kumar 1. Opp. Khuda Baksh Library, Ashok Rajpath, Patna - 4 2. House no. 5A/65, Opp. Mahual Kothi, Alpana Market, Patna Page No.: 2 a) rate = $k[\text{MnO}_4^{2-}][\text{H}_3\text{IO}_6^{2-}]$ b) rate = $k[\text{MnO}_4^{2-}][\text{H}_3\text{IO}_6^{2-}]$ c) rate = $k[\text{MnO}_4^{2-}][\text{H}_3\text{IO}_6]$ d) rate = $k[\text{MnO}_4^{2-}]$ e) rate = $k[\text{MnO}_4^{2-}][\text{H}_3\text{IO}_6^{2-}]^2$ 4.

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PHYSICAL CHEMISTRY By: Shailendra Kumar 1. Opp. Khuda Baksh Library, Ashok Rajpath, Patna - 4 2. House no. 5A/65, Opp. Mahual Kothi, Alpana Market, Patna Page No.: 1 01) Calculate osmotic pressure of a decimolar solution of glucose at 30°C. Ans : 2.488 atm 02) The osmotic pressure of a non volatile solute in C_6H_6 at 25°C is 20.66 Nm⁻² ...

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PHYSICAL CHEMISTRY By: Shailendra Kumar 1. Opp. Khuda Baksh Library, Ashok Rajpath, Patna - 4 2. House no. 5A/65, Opp. Mahual Kothi, Alpana Market, Patna Page No.: 2 (c) $E_a = 35 \text{ kJ/mol}$ $\Delta E = 10 \text{ kJ/mol}$ Ans : Reaction (c) is fastest and reaction (b) is slowest 6) A certain first-order reaction has a rate constant of $1.75 \times 10^{-1} \text{ sec}^{-1}$ at 20°C. What is

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If methyl orange is indicator, then 25 mL of solution requires 25.9 mL of the same HCl for end point. Calculate concentration of each substance in g per litre. [Ans:- 0.1048 g/ litre & 0.1357 g/litre] PHYSICAL CHEMISTRY Page No.: 3 By: Shailendra Kumar

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PHYSICAL CHEMISTRY By: Shailendra Kumar PHYSICAL ... PHYSICAL CHEMISTRY By: Shailendra Kumar 1 Opp Khuda Baksh Library, Ashok Rajpath, Patna - 4 2 House no 5A/65, Opp Mahual Kothi, Alpana Market, Patna Page No: 3 (c) 200 ml of 30 M NaCl is added to 300 ml of 40 M NaCl (d) 200 ml of 20 M BaCl 2 is added to 400 ml of 30 M BaCl 2 ...

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