

Chemistry Chapter 3 Equations Worksheet 1

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3.zinc + hydrochloric acid (zinc chloride + hydrogen. 4.iron + copper(II) sulfate (iron (II) sulfate + copper. 5.copper + sulfuric acid (copper (II) sulfate + hydrogen. 6.potassium dichromate + sulfur + water (sulfur dioxide + potassium hydroxide + chromium (III) oxide.

BALANCING CHEMICAL EQUATIONS - PRACTICE SHEET # 2

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FORMULAS GIVEN Practice Sheet #1 Balance the following equations: 1. $\text{Al} + \text{N}_2 \rightarrow$ 2. $\text{Fe} +$ 3.

$\text{CaCO}_3 \rightarrow$ 4. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$ 5. KI 6. $\text{Pb}(\text{NO}_3)_2 + \text{HCl} \rightarrow$ 7. $\text{BaO}_2 \rightarrow$ 8. $\text{Al} + \text{H}_2\text{SO}_4 \rightarrow$ 9. $\text{CH}_4 + \text{Cl}_2 \rightarrow$ 10. $\text{MgCl}_2 + \text{NaOH} \rightarrow \text{Mg}(\text{OH})_2$ 11. $\text{AgNO}_3 + \text{CuCl}_2 \rightarrow \text{AgCl}$ 12. ZnS 13. Na 14.

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The chemical equation also does not give any information about the speed of the reaction. Some chemical equations and reactions have diverse affect. Students likely find difficulty in balancing chemical equations worksheet. To help you resolve this issue, we have balancing equations worksheet with answers on our main website.

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Worksheet #3: Decomposition Reactions In decomposition reactions, one compound will break down into two or more parts. 1. barium carbonate \rightarrow barium oxide + carbon dioxide $\text{BaCO}_3 \rightarrow \text{BaO} + \text{CO}_2$ 2. magnesium carbonate \rightarrow magnesium oxide + carbon dioxide $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$ 3.

Worksheet #3: Decomposition Reactions

Stoichiometry, part 3; Chapter 5 (Introduction to Solutions and Aqueous Reactions) Solution concentration; Dilutions; Making solutions; Molarity; Titrations; Molecular, Full, and Net-ionic equations; Chapter 6 (Gases) Gas laws; Gas stoichiometry; Chapter 7 (Thermochemistry) Calorimetry worksheet; Calorimetry exercises and answers

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Equation 11.3.2 is identical to Equation 11.3.1 except for the addition of abbreviations in parentheses to indicate the physical state of each species. The abbreviations are (s) for solid, (l) for liquid, (g) for gas, and (aq) for an aqueous solution, a solution of the substance in water.. Consistent with the law of conservation of mass, the numbers of each type of atom are the same on both ...

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